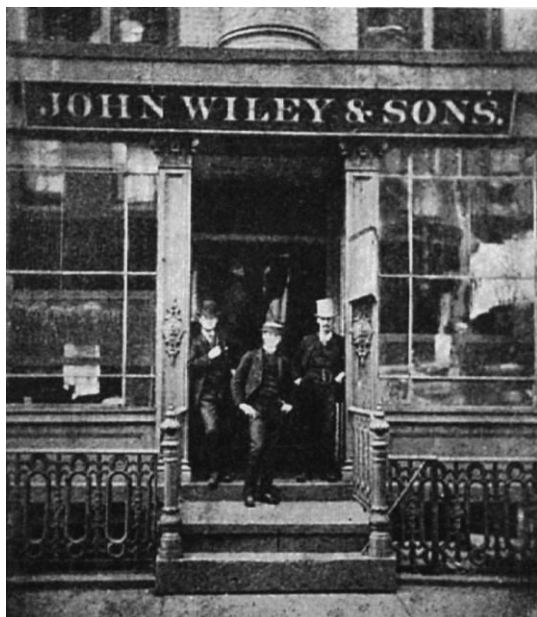


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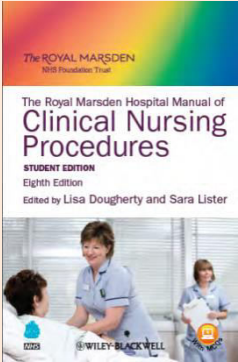


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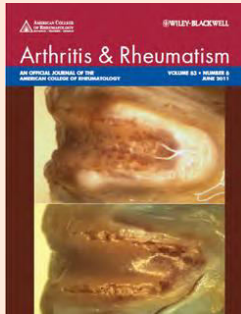



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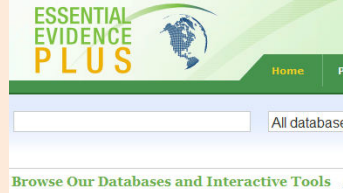

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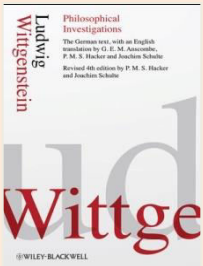
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
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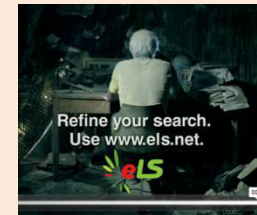
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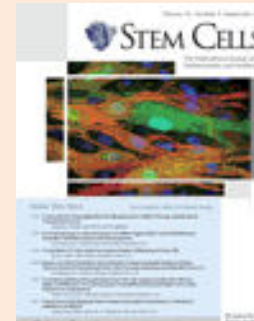
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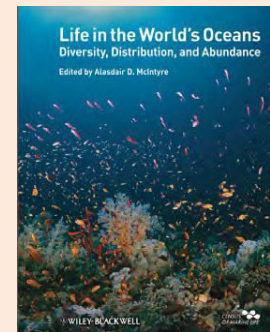

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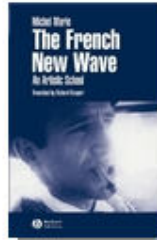
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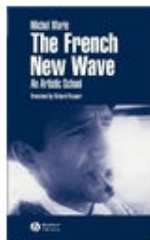
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# Chapter Two

## A Critical Concept

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### A Critical School

**T**HE NEW WAVE DIRECTORS initially and repeatedly denied that they formed a unified movement. For instance, François Truffaut proclaimed: "I see only one common point shared by young directors: we all play pinball while older directors drink scotch and play cards."<sup>1</sup> Despite the many denials, however, this artistic trend was closely tied to a collection of critical concepts held by a fairly coherent group. One of the first criteria for belonging to the New Wave was to have experience as a film critic. These

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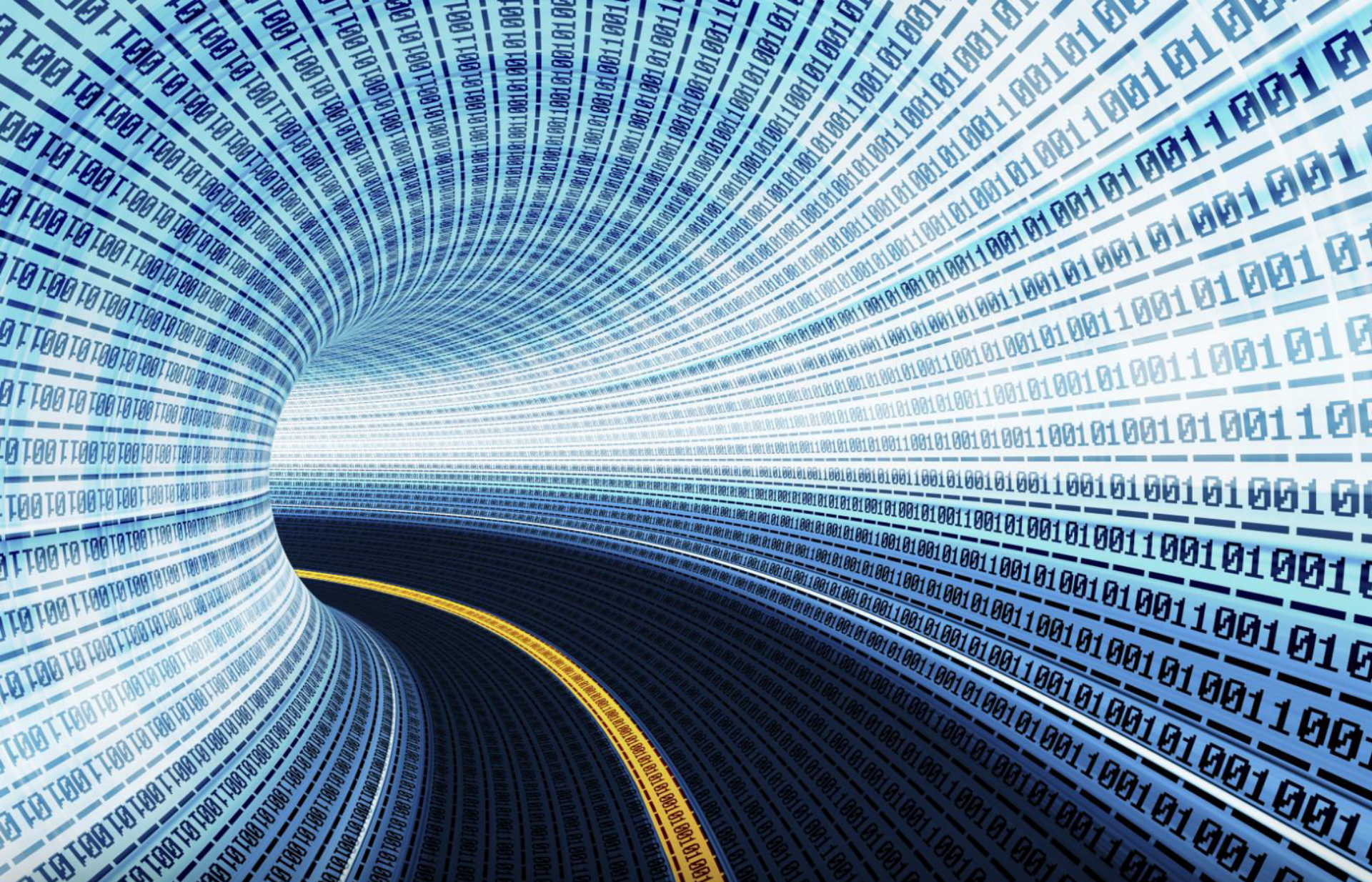
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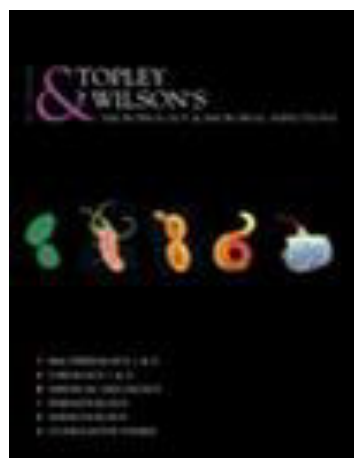
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## 1 Introduction

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Acetic acid [64-19-7],  $\text{CH}_3\text{COOH}$ , *M*, 60.05, is found in dilute solutions in many plant and animal systems. Vinegar, an aqueous solution containing about 4 – 12 % acetic acid, is produced by the fermentation of wine and has been known for more than 5000 years.

The major producers of acetic acid, accounting for ca. 70 % of total worldwide production, are the United States, Western Europe, and Japan. World capacity exceeds  $7 \times 10^6$  t/a [1]. The largest end uses are in the manufacture of vinyl acetate [108-05-4] and acetic anhydride [108-24-7]. Vinyl acetate is used in the production of latex emulsion resins for applications in paints, adhesives, paper coatings, and textile treatment. Acetic anhydride is used in the manufacture of cellulose acetate textile fibers, cigarette filter tow, and cellulose plastics.

## 2 Physical Properties

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Acetic acid is a clear, colorless, corrosive liquid that has a pungent odor and is a dangerous vesicant. It has a  $\text{p}K_a$  of 4.77 [2]. It melts at 16.75 °C [3] and boils at 117.9 °C [4] under 101.3 kPa [5]. It has a pungent vinegarlike odor. The detectable odor is as low as 1 ppm. The acid is combustible with a low flash point of 43 °C. The explosion limits of acetic acid vary from the upper explosion limit (UEL) of 16 % at

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Other impurities include propionic acid and butyric acid. *sec*-Butoxy radicals undergo  $\beta$ -scission to yield propionaldehyde. Oxidation of the aldehyde gives propionic acid. Butyric acid is derived from *n*-butyl radicals.

Higher paraffins are oxidized by similar mechanisms to that of butane. However, the products include shorter chain methyl ketones and difunctional intermediates. These are the results of intramolecular hydrogen abstraction. European countries usually produce acetic acid from naphtha since naphtha is cheap and available. However, the naphtha-based processes yield large amounts of impurities that increase the cost of purification.

**Industrial Operation** [6, 114-116] (Fig. 6). Air or oxygen-enriched air can be used as the oxidant. Multivalent metal ions, such as Mn, Co, Ni, and Cr, are used as catalysts. Some processes, however, are noncatalytic. Reaction conditions are 150 – 200 °C for a range of reaction pressures. The solvent consists of acetic acid, varying in composition below some maximum level appears to be critical.

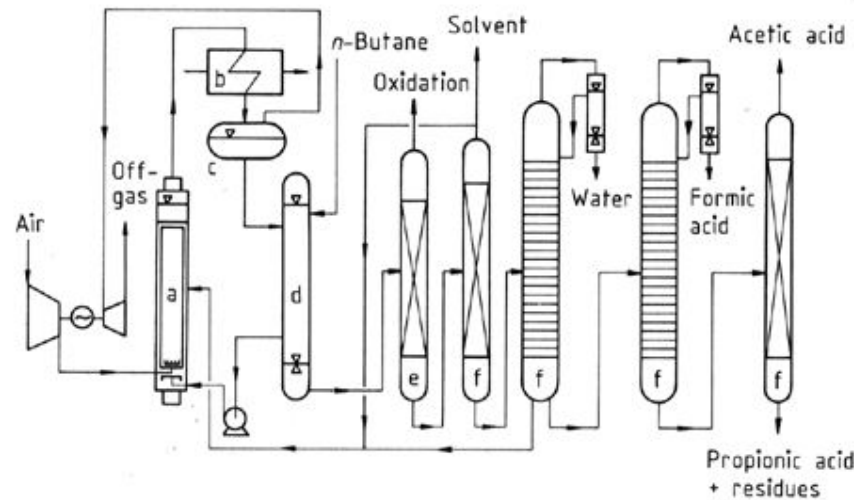


Figure 6. Oxidation of *n*-butane in the liquid phase (Chemische Werke Hüls process)

a) Reactor; b) Air cooler; c) Collector; d) Separation vessel; e) Pressure column; f) Distillation column

The reaction products are separated in hydrocarbon vapor exiting the reactor (a) can be recovered to compress air for use in the reactor. The hydrocarbon vapor exiting the reactor (a) can be recovered to compress air for use in the reactor.

The residues, water, volatile butyrolactone, and other by-products are separated from the acetic acid stream.

The volatile by-products are separated from the acetic acid stream.

The separation of water and formic acid from acetic acid involves several distillations (f). Water removal is the most difficult and costly step of the purification process. It is accomplished by azeotropic distillation with entrainment agents, such as ethers, or by extractive methods.

Formic acid is separated from the resulting anhydrous acetic acid usually by fractionation with an azeotroping agent. The remaining higher boiling acids are separated from acetic acid as a residue stream by fractionation.

In some cases, hydrocarbons such as heptane and isooctane can be added to the reactor to improve separation. These chemicals form a high-boiling azeotrope with formic acid.

In the case of naphtha oxidation, diacids such as succinic acids are isolated for sale [117]. The nonvolatile residue can be burned to

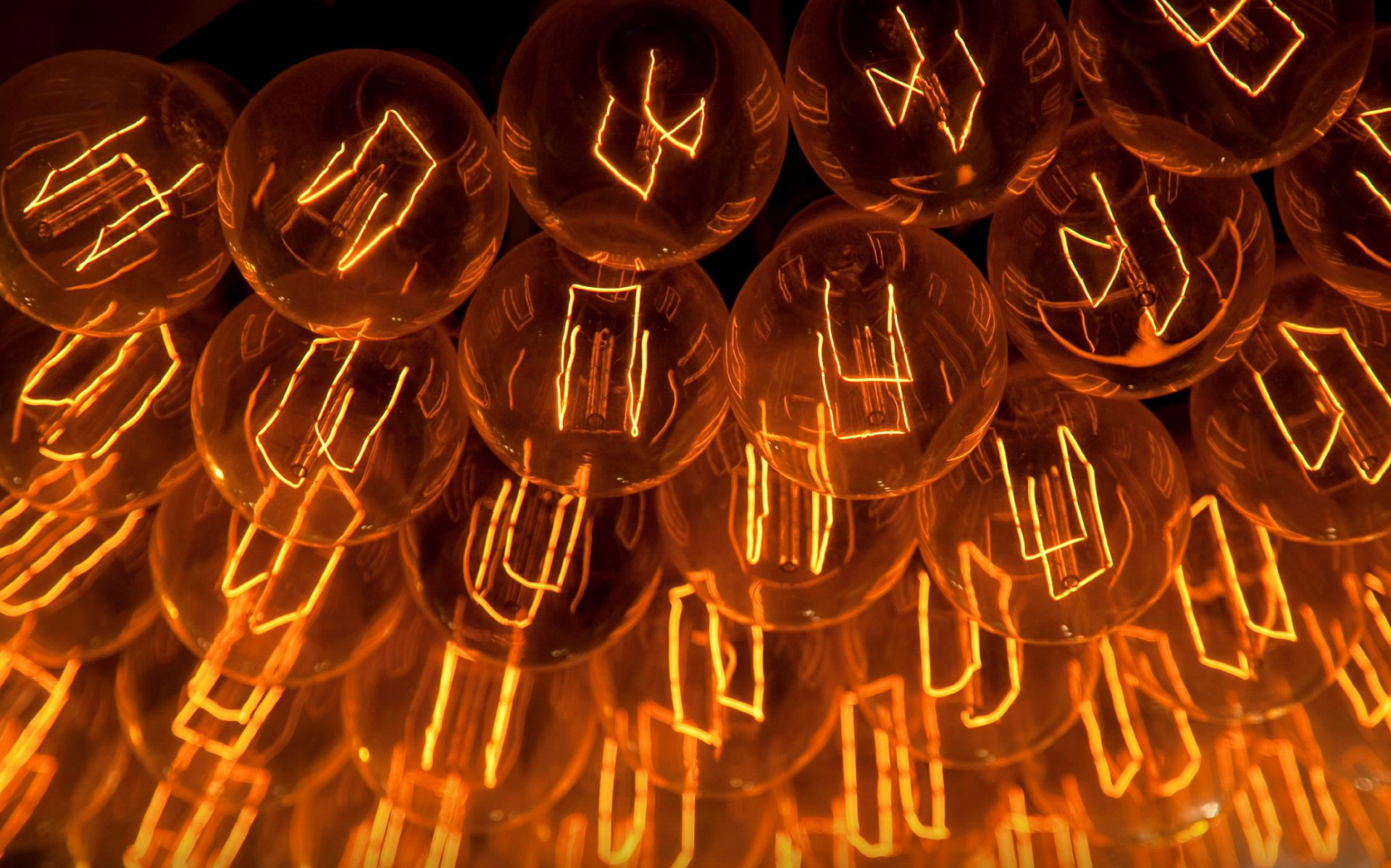
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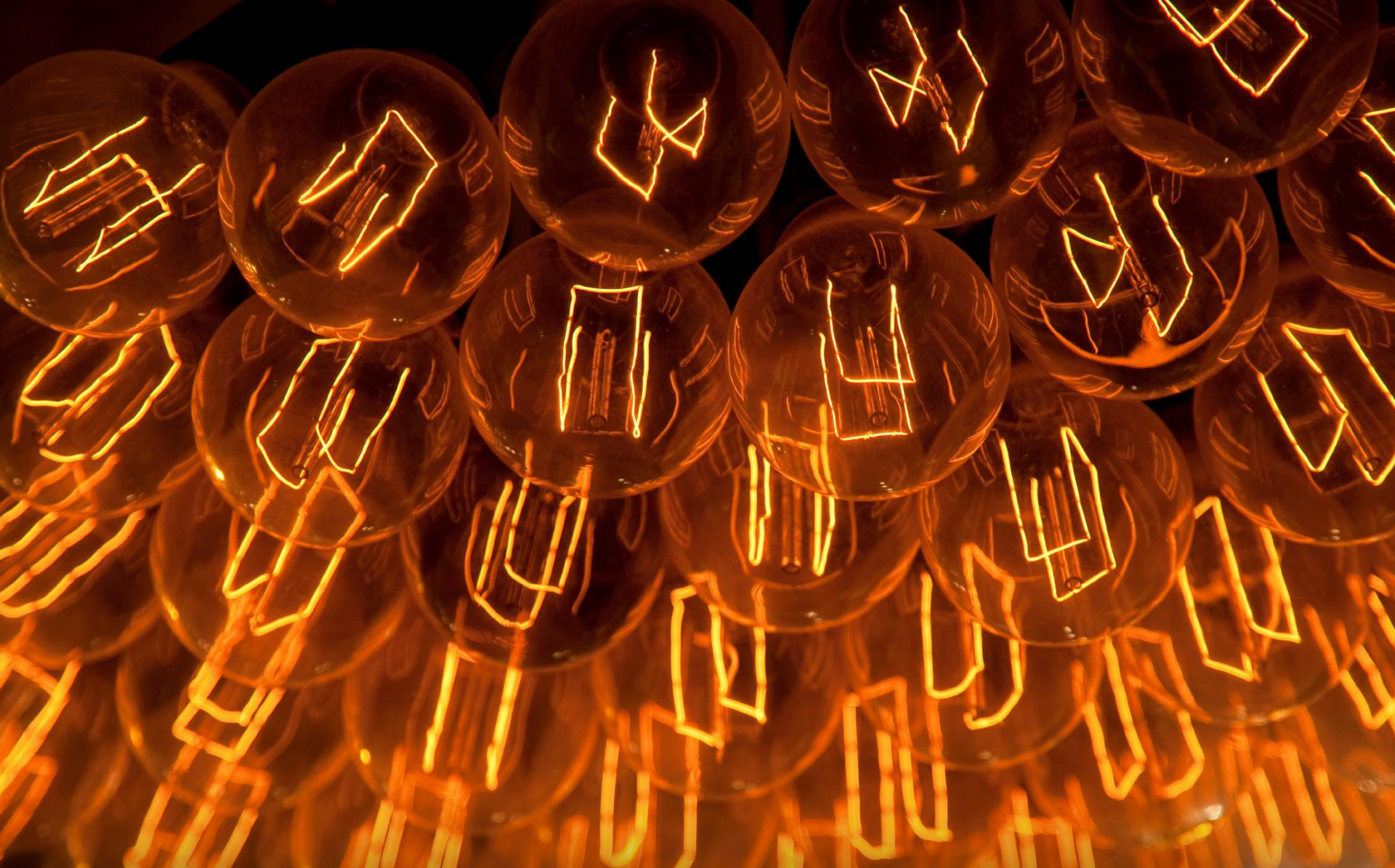
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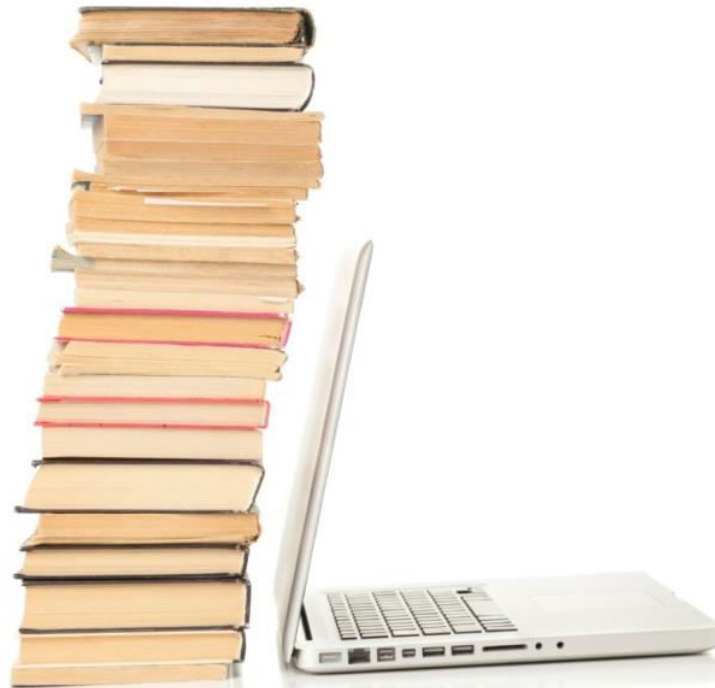
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